Enhancing the Performance of Bi$_2$S$_3$ in Electrocatalytic and Supercapacitor Applications by Controlling Lattice Strain

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Lattice-strained Bi$_2$S$_3$ with 3D hierarchical structures are prepared through a top-down route by a topotactic transformation. High-resolution transmission electron microscopy and X-ray diffraction (XRD) confirm the lattice spacing is expanded by prolonged sulfuration. Performance studies demonstrate that Bi$_2$S$_3$ with the largest lattice expansion (Bi$_2$S$_3$-9.7%, where 9.7% represents the lattice expansion) exhibits a greater electrocatalytic hydrogen evolution reaction (HER) activity compared to Bi$_2$S$_3$ and Bi$_2$S$_3$-3.2%. Density functional theory calculations reveal the expansion of the lattice spacing reduces the bandwidth and upshifts the band center of the Bi 3d orbits, facilitating electron exchange with the S 2p orbits. The resultant intrinsic electronic configuration exhibits favorable H$^+$ adsorption kinetics and a reduced energy barrier for water dissociation in hydrogen evolution. Operando Raman and post-mortem characterizations using XRD and X-ray photoelectron spectroscopy reveal the generation of pseudo-amorphous Bi at the edge of Bi$_2$S$_3$ nanorods of the sample with lattice strain during HER, yielding Bi$_2$S$_3$-9.7%-A. It is worth noting when Bi$_2$S$_3$-9.7%-A is assembled as a positive electrode in an asymmetric supercapacitor, its performance is greatly superior to that of the same device formed using pristine Bi$_2$S$_3$-9.7%. The as-prepared Bi$_2$S$_3$-9.7%-A//activated carbon asymmetric supercapacitor achieves a high specific capacitance of 307.4 F g$^{-1}$ at 1 A g$^{-1}$, exhibiting high retention of 84.1% after 10 000 cycles.

1. Introduction

Bismuth sulfide (Bi$_2$S$_3$) is a layered semiconductor with a direct bandgap of 1.3–1.7 eV$^{[1–3]}$. The bismuth and sulfur atoms are connected by covalent bonds in a 2D layer, and 2D layers are connected by weak van der Waals forces.$^{[4,5]}$ The interlayer spacing of Bi$_2$S$_3$ is 0.32 nm, which is about half that of MoS$_2$.$^{[6]}$ Bi$_2$S$_3$ has aroused great interest due to its high photoconductivity, good environmental compatibility, and abundant natural reserves.$^{[7]}$ However, up till now, Bi$_2$S$_3$ has been used in the field of photoelectrochemistry,$^{[8]}$ photoelectronics,$^{[9]}$ CO$_2$ reduction,$^{[10,11]}$ and nitrogen fixation,$^{[12–14]}$ and there is no research on the application of pure Bi$_2$S$_3$ in electrocatalytic hydrogen evolution reaction (HER)$^{[15,16]}$. Similar to MoS$_2$, due to the bismuth monoatomic layer sandwiched between two sulfur monoatomic layers in the 2D planar structure of Bi$_2$S$_3$, only the unsaturated sulfur edge atoms in Bi$_2$S$_3$ have proved catalytically active in photochemical applications, while the basal plane of the 2D structure is chemically inert.$^{[17–19]}$ This prevents Bi$_2$S$_3$ from contacting reactant molecules and greatly impacts its electrochemical performance.$^{[20]}$

Applying lattice strain is an effective method to modify the internal interatomic distance, potentially increasing the contact with solvent molecules and changing the geometry and electronic structure of the active site, which helps to optimize electrocatalytic activity.$^{[21,22]}$ Due to the complex structure of inorganic metal sulfides and the difficulty of characterization, the lattice strain produced in previous studies of metal sulfides is usually related to structure mismatch, substrates induced or heteroatom substitution, and these factors make it difficult to identify strain-activity correlations.$^{[23–26]}$ However, the unique periodic layered structure and the loosely stacked framework with large open channels of Bi$_2$S$_3$ provide the opportunity of applying lattice strain without introducing the aforementioned factors.

Herein, a top-down route of topological transformation is described that permits the preparation of Bi$_2$S$_3$ with 3D multi-level architectures and different levels of strain. The influence of sulfuration time on lattice strain is determined using high-resolution transmission electron microscopy (HRTEM) and X-ray diffraction (XRD). The HER electrocatalytic activity...
of Bi2S3 samples with different lattice strain is investigated and rationalized with reference to X-ray photoelectron spectroscopy (XPS) and ultraviolet-visible spectroscopy (UV–vis) spectroscopy studies and density functional theory (DFT) calculations.

2. Results and Discussion

2.1. Structural Characterization and Chemical Analysis

2.1.1. Preparation and Lattice Strain Analysis

Bi2S3 hierarchical structures with lattice strain are synthesized via a topotactic transformation method, as illustrated in Figure 1a. The hierarchical BiOCOOH was prepared as a precursor, and it was sulfurized to yield Bi2S3 with similar architectures to the parent material. In this procedure, the crystal structure changes to that of Bi2S3 while the spatial morphology of the precursor is maintained, this is due to ion exchange of S2− with O2− and COOH− and lattice matching of BiOCOOH and Bi2S3. By prolonging the conversion time to Bi2S3, the interlayer interaction is weakened and the interlayer distance expanded, resulting in increased lattice expansion. Bi2S3 with different lattice strains is obtained by adjusting the sulfuration time.

Figure 1b displays a scanning electron microscopy (SEM) image of the hierarchical BiOCOOH assembled from interlaced nanosheets. Figure 1c–e is the SEM and high-resolution TEM (HRTEM) images of the hierarchical BiOCOOH assembled from interlaced nanosheets. Figure 1c is the SEM image of the hierarchical BiOCOOH, Figure 1d is the high-resolution TEM image of the hierarchical BiOCOOH, and Figure 1e is the high-angle annular dark-field (HAADF) scanning transmission electron microscopy image of the hierarchical BiOCOOH. The corresponding selected area electron diffraction (SAED) pattern (Figure 2c) displays unique four-strong ordered electron diffraction spots, similar to single-crystal diffraction spots, indicating these nanorods have high crystallinity and a preferred growth direction of [001]–oriented Bi2S3 nanorods along the two perpendicular [100] and [010] directions. The tendency of the growth of [001]-oriented Bi2S3 nanorods along the two perpendicular [100] and [010] directions is due to the close lattice matching between the a- or b-axis of tetragonal BiOCOOH (a = b = 3.90 Å) and the c-axis of orthorhombic Bi2S3 (c = 3.98 Å), which finally leads to the interwoven Bi2S3 networks. The high-angle annular dark-field scanning transmission electron microscopy (HAADF-STEM) image and the corresponding element mappings (Figure 2d) confirm the as-prepared Bi2S3 only contains Bi and S. The specific surface area of Bi2S3-9.7% was measured by nitrogen adsorption/desorption isotherms. As displayed in Figure 3a (Supporting Information), the specific surface area of the sample is 179.13 m2 g−1, showing a much higher catalytic surface area than that of the commercial Bi2S3.[28] The pore volume (Figure 3b, Supporting Information) of Bi2S3-9.7% calculated by the Barrett–Joyner–Halenda (BJH) method is 0.102 cm3 g−1, and the pores are mainly distributed at less than 1, 2.1, and 3.6 nm, confirming the co-existence of micropores.
and mesopores. The results indicate the 3D hierarchical structure greatly enhanced the absorption ability and pore structure of the sample.

The energy level diagrams of Bi$_2$S$_3$ and Bi$_2$S$_3$-9.7% are derived from the XPS valence band and UV–vis absorption spectra, as shown in Figure 2e,f and Figure S4 (Supporting Information). The valence band maxima of Bi$_2$S$_3$ and Bi$_2$S$_3$-9.7% are 1.01 and 1.84 eV, and the bandgaps are 1.66 and 0.64 eV, respectively. From the energy level diagram, it is seen that the bandgap is decreased by interlayer swelling, indicating the lattice strain enhances the conductivity of the sample.

2.2. Evaluation of the Hydrogen Evolution Performance

The lattice strain alters the local atomic and electronic environment in the Bi$_2$S$_3$ samples, thereby affecting their catalytic activity. The electrocatalytic HER activities were evaluated in 0.5 M H$_2$SO$_4$, 1 M KOH, and 1 M PBS electrolyte using a three-electrode system, all potentials quoted are versus reversible hydrogen electrode (RHE). As shown in Figure 3a–c and Table S1 (Supporting Information) in acidic solution, the required overpotentials for Bi$_2$S$_3$-9.7% are 22, 91, and 127 mV to reach an anodic current density of 10, 50, and 100 mA cm$^{-2}$.
respectively, which are much lower than those of Bi₂S₃ (81, 161, and 220 mV) and Bi₂S₃-3.2% (52, 118, and 163 mV). The fitted value of Tafel slope for Bi₂S₃-9.7% is 72.6 mV dec⁻¹, much smaller than those of Bi₂S₃ (91.4 mV dec⁻¹) and Bi₂S₃-3.2% (77.2 mV dec⁻¹). The observed Tafel slope of Bi₂S₃-9.7% suggests that the mechanism of HER is Volmer (H⁺ + e⁻ → H*) – Heyrovsky (H⁺ + e⁻ + H* → H₂ + H₂O + e⁻). [30,31] Bi₂S₃-9.7% also exhibits the best performance among the samples in both alkaline and neutral electrolytes (Figure 3d–i; Tables S2 and S3, Supporting Information), the overpotentials at 10, 50, and 100 mA cm⁻² are 74, 143, and 175 mV in 1 M KOH, and 96, 209, 269 mV in 1 M PBS, and the Tafel slopes are 73.8 and 87.2 mV dec⁻¹ in alkaline and neutral electrolytes (Figure 3d–i; Tables S2 and S3, Supporting Information), the overpotentials at 10, 50, and 100 mA cm⁻² are 128, 186, 228 mV and 177, 327, 397 mV in 1 M KOH, respectively, and 158, 280, 354 mV and 243, 361, 425 mV in 1 M PBS, respectively. The Tafel slopes of Bi₂S₃ and Bi₂S₃-3.2% are 81.4 and 107.5 mV dec⁻¹ in alkaline electrolyte, and 90.1 and 95.9 mV dec⁻¹ in alkaline electrolyte, respectively (Figure 3d–i; Tables S2 and S3, Supporting Information). The similar values of Tafel slopes for the Bi₂S₃ samples indicate the Volmer–Heyrovsky mechanism across the pH range. The Tafel slopes in acidic media are smaller than those in alkaline and neutral electrolytes because the Volmer step requires a large amount of activation energy in alkaline and neutral solutions, which leads to relatively slower HER reaction kinetics. [32]

In view of the good activity in all pH media, the Bi₂S₃ samples were tested for HER performance in seawater (Figure 4a,b; Table S4, Supporting Information). Bi₂S₃-9.7% exhibits far better catalytic activity than the other samples, showing the lowest overpotentials (η₁₀ = 126 mV, η₅₀ = 221 mV, η₁₀₀ = 303 mV) and the lowest Tafel slope (85.1 mV dec⁻¹), further proving the activity enhancement from lattice strain. The presence of lattice strain also prolongs the electrocatalytic HER stability of Bi₂S₃-9.7% in seawater, and there is no degradation in performance after 60 h (Figure 4c). The N,N-diethyl-p-phenylenediamine (DPD) method is used to detect the possible formation of Cl₂-oxidation products (ClO⁻) during the test (Figure 4c, inset). No color change indicates no production of the hypochlorite that corrodes the catalyst. [33,34] The intrinsic catalytic activity of the Bi₂S₃ samples is evaluated by determining the electrochemically active surface area (ECSA) and mass activity. ECSA is proportional to the double-layer capacitance (Cdl), which is obtained and calculated from cyclic voltammetry (CV) curves, as shown in Figure 4d and Figure S5 (Supporting Information). The Cdl of Bi₂S₃-9.7% is 50.25 mF cm⁻², much higher than those of Bi₂S₃ (42.75 mF cm⁻²) and Bi₂S₃-3.2% (34.25 mF cm⁻²). The mass
activity of Bi$_2$S$_3$-9.7% at an overpotential of 60 mV in freshwater is 80.3 A g$^{-1}$ (Figure 4e), which is ≈ five times and two times higher than that of Bi$_2$S$_3$ and Bi$_2$S$_3$-3.2%, respective, which implies that the lattice strain increases the catalytic activity. In addition, the electrolyte contact angle test demonstrates that the Bi$_2$S$_3$ with the largest lattice swelling (Bi$_2$S$_3$-9.7%) has the highest hydrophilicity (Figure 4f), indicating that the lattice expansion effectively increases the contact between reactants and solvent molecules, leading to a better catalytic performance. The remarkable HER activity of the as-obtained Bi$_2$S$_3$-9.7% is superior to most of the state-of-the-art Bi- and S-based non-noble metal HER electrocatalysts previously reported in all pH electrolytes and seawater (Figure 4g–i; Tables S5–S7, Supporting Information).

2.3. Active Sites for Hydrogen Evolution Catalysis

2.3.1. Interpretation of the Lattice Strain on the Influence of HER Activity

Theoretical simulations were performed to gain further insight into the relationship of the improved electrocatalytic activities and lattice strain. The geometric structures of Bi$_2$S$_3$ samples with different lattice strains are shown in Figure 5a, these structures were built in consideration of generated lattice strain, and the corresponding bulk moduli were calculated, which is inversely proportional to the experimental lattice strain.$^{[35–37]}$ The calculated density of states (DOS) of Bi$_2$S$_3$-9.7% near the Fermi level is higher and the bandgap is smaller compared with those of the other samples (Figure 5b), implying enhanced electron mobility, consistent with the XPS and UV–vis results. The optimized structures and Gibbs free energy diagrams of Bi$_2$S$_3$ samples for HER in acidic solution are shown in Figure 5c. It is observed that the $|\Delta G_{\text{H}}|$ of Bi$_2$S$_3$-9.7% ($-0.28$ eV) is much smaller and closer to the optimal value of zero than those of Bi$_2$S$_3$ ($-0.89$ eV) and Bi$_2$S$_3$-3.2% ($-0.53$ eV), revealing the favorable H$^*$ adsorption after the introduction of lattice strain. The optimized structures and free energy diagrams in alkaline and neutral solutions are displayed in Figure 5d, in which the dissociation of water is pivotal. The energy barrier for water dissociation in the Volmer reaction on the surface of Bi$_2$S$_3$-9.7% is 0.67 eV, which is lower than those of Bi$_2$S$_3$ (1.09 eV) and Bi$_2$S$_3$-3.2% (0.88 eV), and the barriers on the Heyrovsky step for Bi$_2$S$_3$-9.7%, Bi$_2$S$_3$-3.2%, and Bi$_2$S$_3$ are 1.12, 1.45, and 1.63 eV, respectively. The relatively lower barriers of Bi$_2$S$_3$-9.7%
indicate the promoted water dissociation kinetics on the sample with high lattice strain. Furthermore, the influence of lattice strain on Bi 3d orbital configuration was investigated. With the increase of the lattice strain, the Bi 3d band width decreases, and the d-band center shifts positively compared to the Fermi level, implying that the generated strain greatly alters the electronic properties (Figure 5e). A scheme of the changes of the band structures in the samples with lattice strain is shown in Figure 5f. The lattice strain gives rise to a decreased bandwidth and an uplift of the Bi 3d-band center, promoting the electron exchange with the S 2p orbits, leading to greater covalencies in the Bi=S bond and enhancing the binding of hydrogen intermediates with neighboring Bi atoms, which is conducive to the improvement of hydrogen evolution catalytic activities.

2.3.2. Post-Mortem Investigations after HER

To understand the catalytically active sites that produce the extraordinary HER activity of the Bi2S3 catalyst with lattice strain, the microstructure, chemical composition and valence state evolution of Bi2S3-9.7% after the HER test (Bi2S3-9.7%-A) in alkaline electrolyte were studied. The SEM image, TEM image, and corresponding SAED pattern of Bi2S3-9.7% after HER show that the hierarchical structure with interwoven nanorods is maintained, indicating its great structural stability (Figure 6a–c). However, it can be seen from the HRTEM images (Figure 6d,e) that a pseudo-amorphous structure is formed on the outside of the nanorods after HER, and the lattice spacing is reduced to 3.3 Å, corresponding to a lattice expansion of 6.4%, indicating a decrease of lattice strain during the test. The XRD test was performed on the sample after HER (Figure 6f), by comparing with the XRD pattern of Bi2S3-9.7% after HER show that the hierarchical structure with interwoven nanorods is maintained, indicating its great structural stability (Figure 6a–c). However, it can be seen from the HRTEM images (Figure 6d,e) that a pseudo-amorphous structure is formed on the outside of the nanorods after HER, and the lattice spacing is reduced to 3.3 Å, corresponding to a lattice expansion of 6.4%, indicating a decrease of lattice strain during the test. The XRD test was performed on the sample after HER (Figure 6f), by comparing with the XRD pattern of Bi2S3-9.7%, a peak appearing at 30.6° corresponding to Bi, indicating that Bi reduction occurs during the test (JCPDS card No. 51-0765), suggesting that the structure at the edge of the Bi2S3 nanorods is quasi-crystalline Bi metal. Shift to high angle in the XRD pattern of Bi2S3-9.7%-A is also observed, indicating that the lattice expansion is reduced, consistent with the HRTEM result. The blue-shift of the Bi=S band...
in FT-IR reveals that the reduced lattice expansion shortens the bond length of Bi=S (Figure 6g). The fitted Bi 4f spectra of Bi$_2$S$_3$ with different lattice strains before and after HER were compared using XPS (Figure 6h; Figure S6, Supporting Information) and found that the chemical composition of the Bi$_2$S$_3$ sample without lattice strain remained basically unchanged before and after the HER test. With the increase of lattice strain, a new metallic Bi peak appeared in the fitted Bi 4f, which is consistent with the results of HRTEM and XRD, confirming the generation of Bi in the samples with lattice strain after HER.[41,42]

Operando Raman spectroscopy was further performed on the Bi$_2$S$_3$ catalysts to explore the effect of lattice strain on the change of valence state in the real-time HER process (Figure 7). As the voltage alters from 0 to −0.2 V versus RHE, the strain-free catalyst remains stable, and the peaks at 123 and 139 cm$^{-1}$ are attributed to the E$_g$ and A$_1g$ stretching modes of Bi=S bonds, respectively (Figure 7a).[43–45] For the strain-containing Bi$_2$S$_3$, the characteristic vibration of Bi=S band gradually weakened, and a new band appeared at 96 cm$^{-1}$, corresponding to the A$_{2g}$ stretching modes of Bi=Bi bonds (Figure 7b,c).[46,47]

In addition, the self-reconstruction of the samples with high strain is more obvious, the time of this process for Bi$_2$S$_3$-9.7% is shorter compared with that of Bi$_2$S$_3$-3.2%, and the new Bi=Bi bands are also stronger for Bi$_2$S$_3$-9.7%, which shows that the lattice strain is beneficial in the self-reconstruction and production of Bi during HER.

By combining ex situ electron microscopy with Operando Raman spectroscopy, it is concluded that lattice strain plays a key role in triggering surface reconstruction and the reduction of Bi$_2$S$_3$ under applied negative potential, and the in situ generated Bi-metallic phase are the active sites for the HER process.

2.4. Capacitive Performance of Bi$_2$S$_3$-9.7% after HER

2.4.1. Charge Storage Kinetics of Bi$_2$S$_3$-9.7% after HER

Bi$_2$S$_3$ has been widely used as an electrode in supercapacitors due to its low price and high theoretical capacity, however, its supercapacitive performance is limited by the low electrical conductivity and self-aggregation during the test.[48–53]
Bi₂S₃-9.7% catalyst after HER (Bi₂S₃-9.7%-A) demonstrates good structural stability, and its conductivity has been greatly improved due to the introduction of lattice strain and the in situ generation of metallic Bi by reduction during the HER process. In view of the benefits discussed above, the supercapacitive behaviors of the Bi₂S₃-9.7%-A electrode have been investigated in a three-electrode configuration in 3 m KOH electrolyte.

Figure 8a shows the CV curves of Bi₂S₃-9.7% electrodes in 3 m KOH before and after HER testing. The Bi₂S₃-9.7%-A electrode exhibits a higher redox current and a larger enclosed area, indicating its enhanced faradic capacitance that results from increased electroactive sites. The redox peaks reveal the reversible conversion between Bi₂S₃ and Bi₂S₃OH in the electrolyte (Bi₂S₃ + OH⁻ → Bi₂S₃OH + e⁻). For Bi₂S₃-9.7%-A, the additional peak area is attributed to the reversible reaction of Bi and Bi₂O₃ (2Bi +6OH⁻ → Bi₂O₃ + 3H₂O + 6e⁻), the electron transfer number of Bi is significantly higher than that of Bi₂S₃, greatly increasing the capacitance. Figure 8b displays the CV curves of Bi₂S₃-9.7%-A electrode at scan rates of 5–25 mV s⁻¹. As the scan rate increases, the peak current response increases, and the oxidation and reduction peaks shift positive and negative, respectively, which is due to the electrochemical polarization and internal resistance of the electrode.

To further study the enhanced supercapacitive behavior of Bi₂S₃-9.7%-A, the CV curve is analyzed. The power law was applied to scrutinize the diffusion-controlled and surface capacitive contributions of current, which is expressed as \( i_p = a \times v^b \), where \( i_p \) is the peak current (A cm⁻²), \( v \) is the scan rate (V s⁻¹), and \( a \) and \( b \) are the adjustable parameters. A \( b \)-value of 0.5 indicates a diffusion-controlled process, whereas a \( b \)-value of 1 suggests surface capacitive behavior. The \( b \)-values of cathodic and anodic peaks calculated from the slopes of the linear fit of log(\( i_p \)) versus log(\( v \)) are 0.60 and 0.53, respectively (Figure 8c), suggesting the current response of Bi₂S₃-9.7%-A electrode is dominated by diffusion-controlled behavior. Figures S7 and S8 (Supporting Information) show the linear fit of \( v^{1/2} \) versus \( i_p^{1/2} \) for Bi₂S₃-9.7%-A electrode. To quantify the relative contribution of the diffusion-controlled...
and surface capacitive mechanisms, the power law is modified as

\[ i_p = k_1 \nu + k_2 \nu^{1/2} \]

where it is changed as \( i_p/\nu^{1/2} = k_1 \nu^{1/2} + k_2 \), where \( i_p \) is the peak current (A cm\(^{-2}\)), \( \nu \) is the scan rate (V s\(^{-1}\)), \( k_1 \) and \( k_2 \) are current contributions from surface capacitive and diffusion-controlled behavior, respectively.\(^{[59–63]}\) The relative current contributions obtained according to the above equations are displayed in Figure 8d. It is observed that at a low scan rate of 5 mV s\(^{-1}\), the diffusion-controlled contribution (87.4%) is much higher than the capacitive current contribution (12.6%). With an increase in the scan rate, the surface capacitive contribution of the fast kinetics process gradually increases, reaching 28.3% at a high scan rate of 25 mV s\(^{-1}\), while it is still lower than the current contribution of the diffusion-controlled process (71.7%). This indicates that the charge storage capacity of the Bi\(_2\)S\(_3\)-9.7%-A electrode is mainly derived from faradaic behavior.

2.4.2. Capacitive Performance of Bi\(_2\)S\(_3\)-9.7%-A//AC Asymmetric Supercapacitor

Encouraged by the excellent electrochemical capacitive performance of Bi\(_2\)S\(_3\)-9.7%-A, an asymmetric supercapacitor (ASC) was assembled using Bi\(_2\)S\(_3\)-9.7%-A as the positive electrode and activated carbon (AC) as the negative electrode, as shown in Figure 8e. Figure 8f displays the CV curves of the Bi\(_2\)S\(_3\)-9.7%-A//AC ASC measured at scan rates between 10 and 50 mV s\(^{-1}\) across a potential window of 0–1.5 V. It is observed that the CV curves of Bi\(_2\)S\(_3\)-9.7%-A//AC ASC exhibit ideal capacitive and faradaic behaviors as the scan rate increases, which shows its good electrochemical reversibility and the benefits of the combination of double-layer capacitive and faradaic mechanisms.

Figure 8g displays the galvanostatic charge–discharge (GCD) curves performed at different current densities of 1–32 A g\(^{-1}\). Bi\(_2\)S\(_3\)-9.7%-A//AC ASC exhibits high gravimetric capacitance of 307 and 133 F g\(^{-1}\) at current densities of 1 and 30 A g\(^{-1}\), superb capacitance and rate capability, much better than that of Bi\(_2\)S\(_3\)-9.7%-A//AC ASC (Figure 8h; Figure S9, Supporting Information). This demonstrates that the in situ generated Bi-metal on the fringe of the Bi\(_2\)S\(_3\) nanorods greatly enhances the charge storage ability. The comparison of energy and power density with data previously reported is presented in the Ragone plot (Figure 8i). Bi\(_2\)S\(_3\)-9.7%-A//AC ASC reaches an excellent energy density of 109 Wh kg\(^{-1}\) at a power density of 984 W kg\(^{-1}\), and even maintains 47 Wh kg\(^{-1}\) at a high-power density of 24.4 kW kg\(^{-1}\). The energy storage capacities of Bi\(_2\)S\(_3\)-9.7%-A//AC ASC are much higher than those of the reported Bi\(_2\)S\(_3\)-based ASC devices (Figure 8i,j, Table S8, Supporting Information). The EIS curves and equivalent circuit of Bi\(_2\)S\(_3\)-9.7%-A//AC ASC are shown in Figure 8k, suggesting a low charge transfer resistance (\(R_{ct}\)), which indicates that the device has good charge transfer kinetics.\(^{[52]}\) The long-term cycling performance evaluated at 1 A g\(^{-1}\) shows the device has good durability with a capacitance retention of 84.1% after 10 000 cycles.

Figure 7. Active sites for hydrogen evolution catalysis. Operando Raman spectra and the corresponding contour plot of a) Bi\(_2\)S\(_3\), b) Bi\(_2\)S\(_3\)-3.2%, and c) Bi\(_2\)S\(_3\)-9.7% obtained from the voltage changing from 0 to −0.2 V (vs RHE).
3. Conclusions

In summary, a novel method for the preparation of 3D hierarchical Bi$_2$S$_3$ architectures with tunable lattice strain is reported. The lattice spacing is confirmed to be expanded with the prolongation of sulfuration. The results of XPS valence band spectra and UV–vis show that the bandgap decreases as the lattice strain increases. The as-prepared Bi$_2$S$_3$ electrocatalyst with 9.7%-lattice strain exhibits the best HER performance among the Bi$_2$S$_3$ samples, exhibiting 22 and 91, 74 and 143, 96 and 209, 126 and 221 mV in acidic, alkaline, neutral electrolytes and simulated seawater, with excellent stability for 60 h test in seawater.

Figure 8. Capacitive Performance of Bi$_2$S$_3$-9.7% after HER. a) CV curves of Bi$_2$S$_3$-9.7% before and after HER in alkaline solution. b) CV curves of Bi$_2$S$_3$-9.7% after HER measured at different scan rates of 5–25 mV s$^{-1}$. c) The power–law dependence of cathodic and anodic peak currents of Bi$_2$S$_3$-9.7% after HER from CV curves. d) Histograms of the capacitive and diffusion-controlled battery-type current contribution of Bi$_2$S$_3$-9.7% after HER at different scan rates. e) Schematic illustration of the assembled Bi$_2$S$_3$-9.7%-A//AC ASC. f) CV, and g) GCD curves of Bi$_2$S$_3$-9.7%-A//AC ASC at various scan rates and current densities. h) Calculated specific capacitance for Bi$_2$S$_3$-9.7%-A//AC ASC and Bi$_2$S$_3$-9.7%-A//AC ASC at various current densities. i) Ragone plot of Bi$_2$S$_3$-9.7%-A//AC ASC. j) Comparisons of the capacitive performance with some recently reported Bi-based catalysts in alkaline electrolyte. k) Cycling stability of the assembled Bi$_2$S$_3$-9.7%-A//AC ASC at the constant current density of 1 A g$^{-1}$. (Inset: Nyquist plots and the corresponding equivalent circuit of Bi$_2$S$_3$-9.7%-A//AC ASC).
DFT simulations reveal that the lattice strain reduces the bandwidth and upshifts the band center of Bi 3d orbits, which promotes the electron exchange with the S 2p orbits, favoring the adsorption of H* kinetics and reducing the energy barrier of water dissociation for HER activities. Post-mortem characterizations reveal the formation of pseudo-amorphous Bi on the boundary of Bi2S3-9.7% nanorods after HER, which significantly enhances its capacitance. Bi2S3-9.7%-A is constructed as an electrode in the ASC, and the as-prepared Bi2S3-9.7%-A//AC ASC displays both high energy and power densities in a wide specific capacitance range and shows good retention of 84.1% after 10,000 cycles at 1 A g−1. This work provides valuable guidance for lattice strain regulation of 2D layered sulfide materials for clean energy conversion and storage applications.

Supporting Information
Supporting Information is available from the Wiley Online Library or from the author.

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Conflict of Interest
The authors declare no conflict of interest.

Data Availability Statement
The data that support the findings of this study are available in the supplementary material of this article.

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